

What is combustion?



- A chemical reaction in which a fuel reacts quickly with oxygen gas
- Exothermic
- Products are CO₂and H₂O andenergy

Combustion of Hydrocarbons

- ☐ Hydrocarbons are molecular compounds made up of hydrogen and carbon (who'd a thunk?)
- □ The combustion of hydrocarbons fuels powers our cars, warms our homes, and generates electricity
- Most hydrocarbons are from fossil fuels

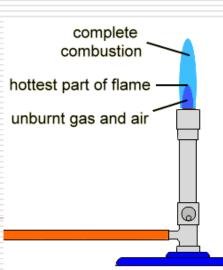


May be complete or incomplete, depending on oxygen supply

- Complete combustion gives the same two products all the time:
 - carbon dioxide and water
- Results in an 'invisible' blue flame







Combustion of Hydrocarbons

Word equation for complete combustion:

Hydrocarbon + oxygen — → carbon dioxide + water

□ General equation:

$$C_xH_y + O_2 \longrightarrow CO_2 + H_2O + energy$$

Combustion of Hydrocarbons

Natural gas is mostly comprised of methane, CH₄

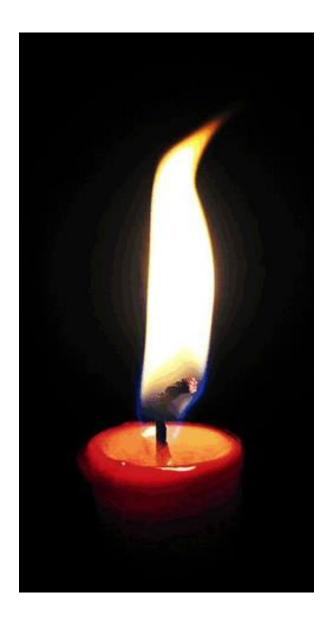
$$CH_4(g) + 2 O_2(g) \longrightarrow CO_2(g) + 2 H_2O(g) + energy$$

During complete combustion, the fuel burns cleanly, leaving no soot or other residue

Incomplete Combustion

- Results when oxygen supply is limited
- As well as carbon dioxide and water, products also include carbon monoxide and soot (carbon)
- Signs of incomplete combustion are an orange, flickering flame, and soot.

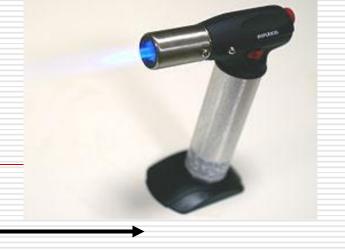
Ex. Portable stoves using butane gas







Incomplete Combustion of Butane

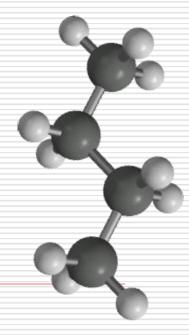


$$C_4H_{10(g)} + O_{2(g)}$$

$$CO_{2(g)} + H_2O_{(g)} + CO_{(g)} + C_{(s)}$$

$$C_4H_{10(g)} + 5O_{2(g)}$$

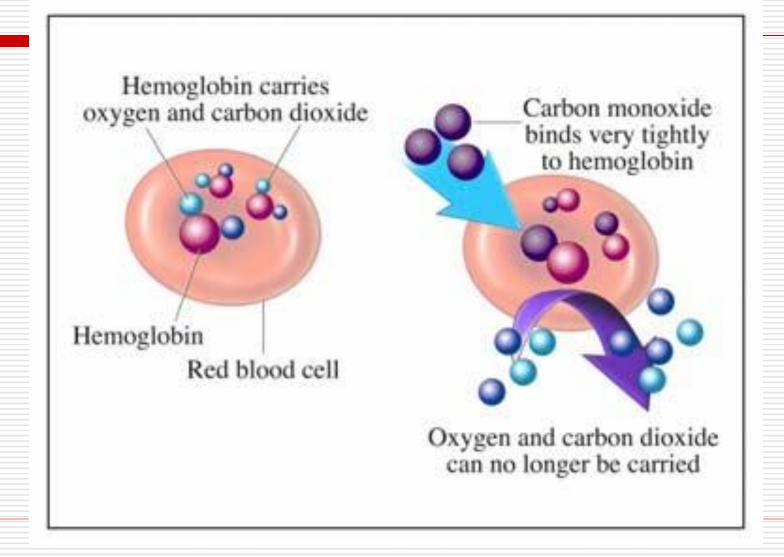
$$2 CO_{2(g)} + 5 H_2O_{(g)} + CO_{(g)} + C_{(s)}$$



Carbon Monoxide

- ☐ Highly toxic
- Can build up in homes with poorly ventilated furnaces or by the combustion of a fuel in an enclosed space
- Symptoms of CO poisoning include headache, dizziness, nausea and difficulty breathing – CAN BE FATAL!

Carbon Monoxide



Soot

- Particles of carbon
- Causes pollution and wastes energy
- Common in older, poorly maintained

fireplaces



Other fuel sources: Combustion of Hydrogen

- Hydrogen can be used as a fuel source
- Benefit: products are water vapour and energy
- Drawbacks: costly to produce & store & transport
- Liability: hydrogen is explosive

$$2H_2(g) + O_2(g) \longrightarrow 2H_2O(g) + energy$$

Get fired up...

- □ Pg 251 # 2 − 3, 7
- CombustionReactions worksheet

